

Name:	Section:	Date:
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Molar Volume of a Gas

Chemical reaction (*with all states of matter*)

Data

Mass of Mg used	_____	g
Volume of gas collected	_____	mL
Room Temperature	_____	°C
Barometric pressure	_____	kPa

Results

Moles of Mg used	_____	mol
Moles of H ₂ produced	_____	mol
Water vapor pressure (from table 1)	_____	kPa
Pressure of H ₂	_____	kPa
Lab value $\bar{V}(\text{H}_2)$ at room conditions	_____	$\frac{\text{L}}{\text{mol}}$
Lab value $\bar{V}(\text{H}_2)$ calculated at STP	_____	$\frac{\text{L}}{\text{mol}}$
Literature value $\bar{V}(\text{H}_2)$ at STP	22.42	$\frac{\text{L}}{\text{mol}}$
%error	_____	%