Name: _____ Date: _____

section: _____

General Chemistry: 202-NYA-05
Solution Stoichiometry (Data and results)

Balanced chemical equation (with the state of matter)

Table of Data (to be filled in ink)

Mass of empty weighing dish		g
Mass of weighing dish + $CaCl_2 \cdot 2H_2O(s)$:		g
Volume of the volumetric flask used ($CaCl_2$):	100.0	mL
$Na_2CO_3(aq)$ concentration (burette)		mol/L
Volume of $Na_2CO_3(aq)$ added (burette)		mL
Volume of $CaCl_2(aq)$ added (pipette)	10.00	mL
Mass of the filter paper + watch glass		g
Mass of the filter paper + watch glass + dry $CaCO_3(s)$		g

Table of Results (to be completed at home)

Mass of $CaCl_2 \cdot 2H_2O(s)$ used for the initial 100 mL solution		g
$CaCl_2(aq)$ concentration (100.0 mL volumetric flask)		mol/L
Moles of $CaCl_2$ added (pipette)		mol
Moles of Na_2CO_3 added (burette)		mol
Limiting reactant (circle one): $CaCl_2$ or Na_2CO_3		
Mass of excess reactant remaining after the reaction		g
Mass of dry $CaCO_3(s)$ recovered		g
Theoretical yield of $CaCO_3(s)$ (mass from calculations)		g
%yield of the reaction		%