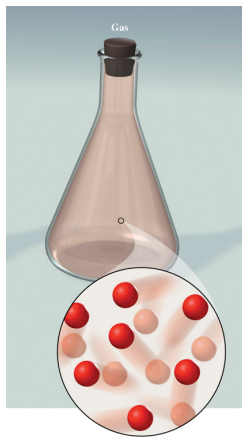
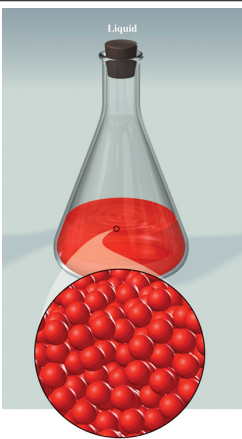


Physical states / Quantitative aspects

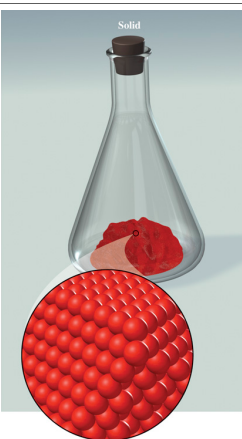
GAS

Intermolecular forces \ll kinetic energy of the molecule.	
Large interparticle distances (random movements).	
Highly compressible.	
Low density (g/L).	
Assume the volume and the shape of the container.	

LIQUID

Intermolecular forces \leq kinetic energy of the molecule.	
Particles in "virtual" contact (random movements).	
Compress only very slightly.	
High density (g/mL).	
Assume the shape of the container with a definite volume.	

SOLID

Intermolecular forces \gg kinetic energy of the molecule.	
Particle close together in a fixed position with/without order.	
Less compressible than liquid.	
High density (g/mL).	
Assume its own volume and shape.	