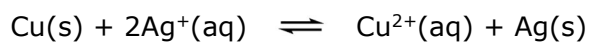
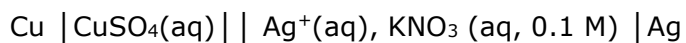
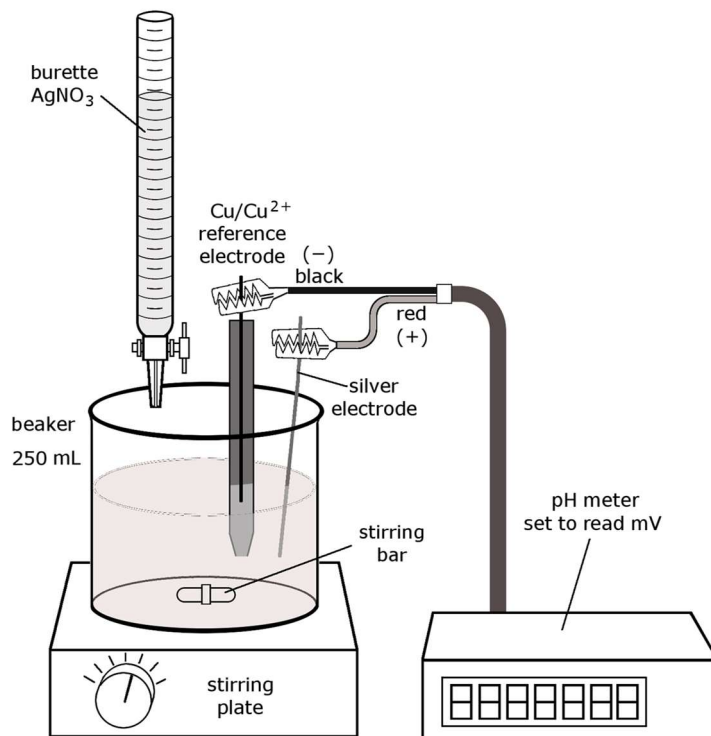
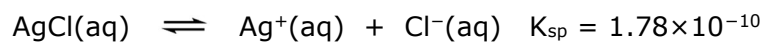
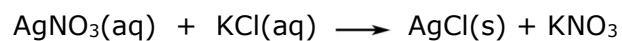


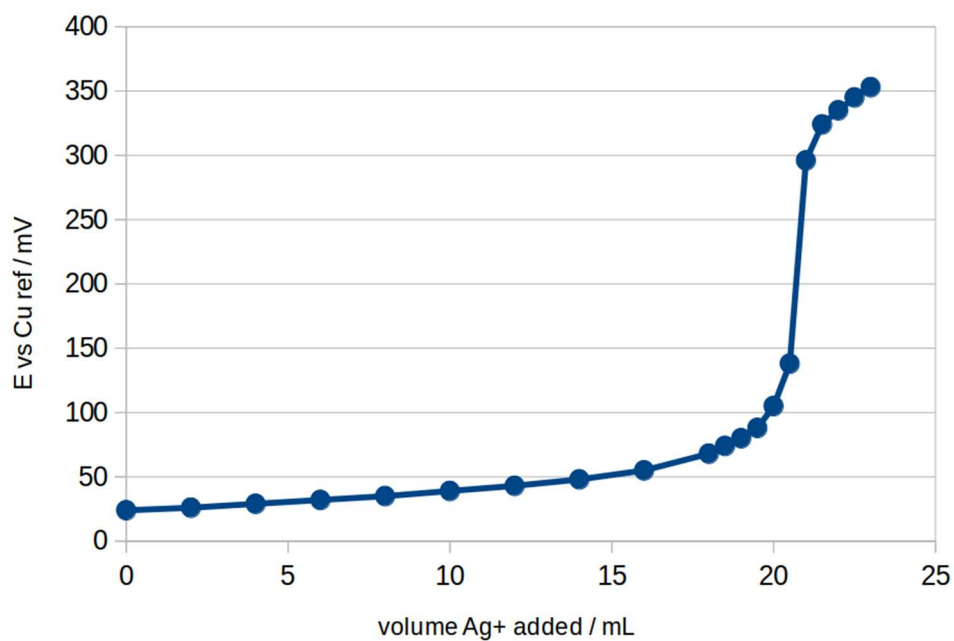
Exp. 8: Potentiometric titration of Cl^-



$$E = (E_{\text{Ag}/\text{Ag}^+}^{\circ} - E_{\text{Cu}/\text{Cu}^{2+}}^{\circ}) - \frac{0.0591\text{V}}{2} \log \frac{[\text{Cu}^{2+}(\text{aq}, \text{ref})]}{[\text{Ag}^+]^2}$$

Cu | CuSO₄(aq) || Ag⁺(aq), KNO₃ (aq, 0.1 M) | Ag

(Actual lab value, Electrochem F2021)



$$E_{cell} = (E_{Ag/Ag^+}^{\circ} - E_{Cu/Cu^{2+}}^{\circ}) - \frac{0.0591V}{2} \log \frac{[Cu^{2+}(aq, ref)]}{[Ag^+]^2}$$

